

CAMBRIDGE INTERNATIONAL EXAMINATIONS

Cambridge International General Certificate of Secondary Education

MARK SCHEME for the May/June 2015 series**0620 CHEMISTRY****0620/51**

Paper 5 (Practical), maximum raw mark 40

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Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Abbreviations used in the Mark Scheme

- ; separates marking points
- / separates alternatives within a marking point
- **OR** gives alternative marking point
- **R** reject
- **I** ignore mark as if this material was not present
- **A** accept (a less than ideal answer which should be marked correct)
- **COND** indicates mark is conditional on previous marking point
- owtte or words to that effect (accept other ways of expressing the same idea)
- max indicates the maximum number of marks that can be awarded
- ecf credit a correct statement that follows a previous wrong response
- () the word/phrase in brackets is not required, but sets the context
- ora or reverse argument

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Question	Answer	Marks	Additional Guidance
1(a)	initial temperature box completed; other temperature boxes completed; trend in temperatures is comparable to Supervisor's results;	3	Temperatures should increase to a maximum then decrease
1(b)	initial temperature box completed; other temperature boxes completed; trend in temperatures is comparable to Supervisor's results;	3	Temperatures should increase to a maximum then decrease
1(c)	all 18 points plotted within half a small square = 3 marks 17 points plotted within half a small square = 2 marks 16 points plotted within half a small square = 1 mark; best fit smooth line/ intersecting straight lines; labels;	5	
1(d)	value read from graph; indication clearly shown;	2	
1(e)	exothermic;	1	
1(f)	to remove traces of acid A/clean; to remove water;	2	
1(g)(i)	experiment 2/ acid B;	1	
1(g)(ii)	acid B is stronger/ dibasic/ has a lower pH/ more acidic;	1	I more reactive/ more concentrated
1(h)	heat losses/ using a measuring cylinder/ thermometer/ cup not washed; insulate/ use burette/ digital thermometer/ new cup;	2	I repeat and average

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Question	Answer	Marks	Additional Guidance
2(a)	white (crystals);	1	A colourless R precipitate
2(b)	melts / liquefies / bubbles / dissolves; steam / condensation / drops of liquid; pH 7–14;	3	A reference to smell I sublimation A colour: green / blue / purple I ammonia
2(c)	white; precipitate; dissolves / clears; pungent gas; pH paper green / blue / purple, pH > 7;	5	
2(d)	white; precipitate;	2	
2(e)	no reaction / no change / no precipitate / nothing;	1	
2(f)	white; precipitate;	2	
2(g)	alkaline gas / ammonia; hydrated / water;	2	
2(h)	not a halide / not a named halide;	1	
2(i)	ammonium / NH_4^+ ; aluminium / Al^{3+} ; sulfate / SO_4^{2-} ;	3	